

Molecular excitations: a new way to detect Dark matter

J.Va'vra

SLAC, Stanford University, CA94309, U.S.A.
e-mail: jjv@slac.stanford.edu

Received February 3, 2014

ABSTRACT

We believe that the Dark Matter (DM) search should be expanded into the domain of detectors sensitive to molecular excitations, and so that we should create detectors which are more sensitive to collisions with very light WIMPs. In this paper we investigate in detail di-atomic molecules, such as Fused Silica material with large OH-molecule content, and water molecules. Presently, we do not have suitable IR detectors to observe single photons, however some OH-molecular excitations extend to visible and UV wavelengths, and can be measured by Bi-alkali photocathodes. There are many other chemical substances with di-atomic molecules, or more complex oil molecules, which could be investigated also. This idea invites searches in experiments having large target volumes of such materials coupled to a large array of single-photon detectors with Bi-alkali photocathodes.

Key words. DAMA experiment, Dark Matter search

1. Introduction

A recent paper [1] discussed the possibility that the DAMA experiment [2] observes the effect of a light WIMP striking a hydrogen atom and creating a proton recoil, which then excites the NaI(Tl) crystal in which it is located. The target hydrogen atom is present as OH or H₂O contamination. The "classical" scintillation mechanism for atomic level excitation in NaI(Tl) is shown schematically in Fig. 1. The primary energy deposit has to exceed the band gap of NaI, which is 3.9 eV; thallium is added to lower the emission energy.

The present paper considers another, even more sensitive, method to detect Dark Matter via molecular excitations. A very small energy-deposit, at a level of 2-3 eV, is needed to excite such vibrations; this is a factor of 2 lower in energy than the classical scintillation mechanism. Many di-atomic, or even more complex, molecules will vibrate when hit by a neutron or a WIMP, and some small fraction of such single photon excitations can be detected by a PMT working in the Bi-alkali regime. Most of the emitted energy is in the IR wavelength region, or is in the form of heat. A preliminary version of these ideas was presented in Ref. [3].

Di-atomic molecules are ideal for such studies, since they are simpler to understand, and have been studied extensively by chemists [4]. The energy levels of di-atomic molecules can be

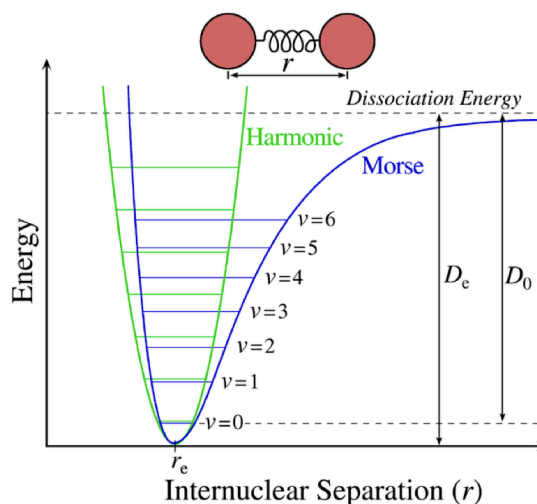


Fig. 2. Energy levels of diatomic molecules for a harmonic and an anharmonic oscillator, where the latter follows the more complex Morse potential; ν is the vibrational quantum number ($\nu = 0, 1, 2, \dots$). The $\nu = 0$ level is the vibrational ground state. Unlike the energy levels of the harmonic oscillator potential, which are evenly spaced by $\hbar\omega$, the Morse potential level spacing decreases as the energy approaches the dissociation energy D_e [5].

Table 1. A simple calculation of the transition wavelength for several frequency overtones of the OH-radicals. The last two modes correspond to visible wavelengths. Higher modes can reach the UV regime [6].

| OH-band identity | Transition | Calculated wavelength [nm] |
|------------------|-------------------|----------------------------|
| ν_1 | $0 \rightarrow 1$ | 2803 |
| $2\nu_1$ | $0 \rightarrow 2$ | 1436 |
| $3\nu_1$ | $0 \rightarrow 3$ | 980 |
| $4\nu_1$ | $0 \rightarrow 4$ | 755 |
| $5\nu_1$ | $0 \rightarrow 5$ | 619.5 |

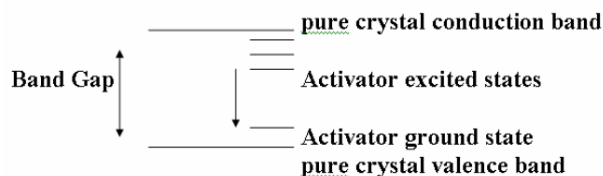


Fig. 1. Energy levels for an activated NaI(Tl) crystal, where Tl is the so-called activator. The Tl-levels lie somewhat below the conduction band so that the radiated photons are emitted with less energy than the band gap of the pure crystal.

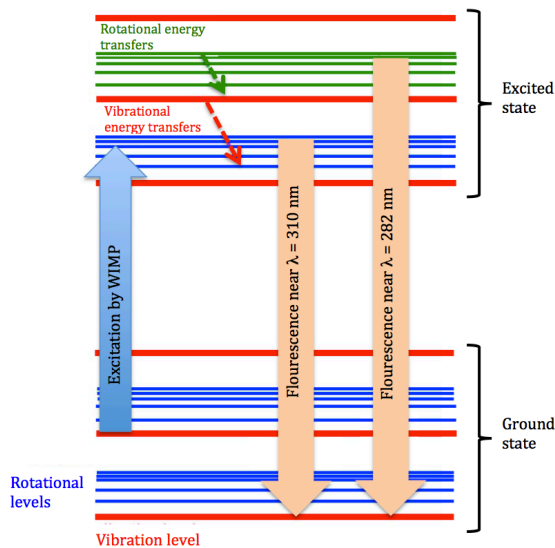


Fig. 3. A schematic picture of WIMP-scattering excitation of energy levels in the OH-molecule, with subsequent de-excitation via fluorescent photons at either ~ 282 nm or ~ 310 nm.

calculated using the Schrodinger equation description of a harmonic oscillator, which has been described in many textbooks - *e.g.* see Refs. [5] and [6]. Figure 2 and Table 1 provide examples of energy levels in a typical di-atomic molecule, for a perfect harmonic oscillator, and for an anharmonic oscillator following the Morse potential [5]. The most significant difference between the two potentials is a transition selection rule. The harmonic oscillator allows only transitions obeying $\Delta\nu = \pm 1$. In contrast, the anharmonic oscillator allows the transitions $\Delta\nu = \pm 1, \pm 2, \pm 3$, etc. These transitions, $\nu_1, 2\nu_1, 3\nu_1$, etc., are called overtones. We see that although the fundamental mode corresponds to an IR wavelength, the higher overtones correspond to visible or even UV wavelengths. This can provide a path to detection using Bialkali photocathode-based PMTs; single-photon IR-sensitive detectors are more difficult to utilize at present for this purpose.

We note that the OH-molecule has been studied extensively by many chemists using laser-induced fluorescence. An example of such a fluorescence measurement is the OH-molecule excitation by a 282 nm dye laser, followed by the observation of the 310 nm wavelength with a PMT with a notch filter [7]. This method was used to determine traces of OH-radicals in the atmosphere [8]. Figure 3 shows schematically how WIMP scattering could excite energy levels in the OH-molecule, with a subsequent de-excitation via fluorescent photons at either ~ 282 nm or ~ 310 nm.

However, the picture outlined above is too simplistic, since the actual energy levels can be affected by the substance in which the OH-radical is located. In this paper we discuss the example of Fused Silica loaded with OH-molecules. Fused Silica is a good example because it has been investigated previously in great detail. It is an ideal material in which to study the effects of the OH-molecular presence, because the basic material is very pure. In addition, Fused Silica can be prepared in the form of long fibers, which allows the accurate measurement of absorption resonances by using a monochromator. The next chapter will deal with this material in more detail.

Although Fused Silica is a well-understood material, pure water is even better understood. We have looked into its molecular vibration response also.

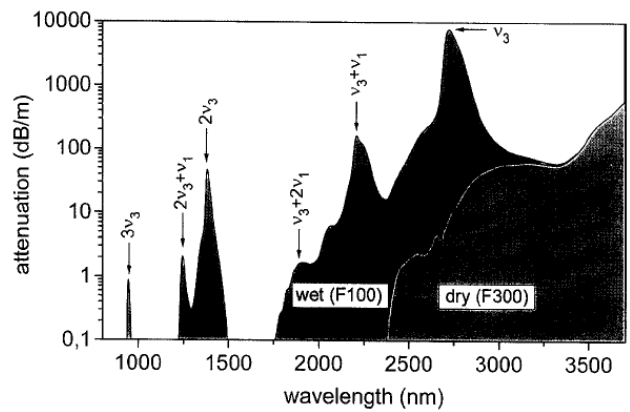


Fig. 4. Attenuation spectra of wet (F100) and dry (F300) Fused Silica short-bulk samples. Each sample is too short (20 cm) to be sensitive to higher harmonic modes [9].

2. OH-molecular vibrations in Fused Silica

Reference [9] investigated IR absorption in wet and dry Fused Silica. Wet Fused Silica contains a large quantity of OH-molecules. In contrast, the dry Fused Silica has almost no OH-content and a minimum of other impurities, so that it is useful for fiber applications operating in the IR regime. The OH content is due to the presence of some moisture or hydrogen in the manufacturing process, which results in hydroxyl groups chemically bonded to the silica molecular network (SiOH). This hydroxyl (OH) affects the optical properties of silica due to the fundamental OH-absorption band with $\nu_3 = 2.72 \mu\text{m}$, the corresponding overtones, and the combination modes with the SiO_4 tetrahedron vibration. The OH content of the dry Fused Silica was 0.2 ppm (designated as F300), and the OH-content of the wet Fused Silica was about 700 ppm (designated as F100). Figure 4 shows the absorption measurements for both types of Fused Silica in short-bulk samples. The spectrum of the wet silica is affected by the fundamental mode ν_3 , the overtones $2\nu_3, 3\nu_3$, and a few combination modes. Higher overtones, $n\nu_3$, were not observed because the sample was only 20 cm in length. However, the fundamental ν_3 frequency was observed in the wet sample. The dry sample does not exhibit major absorption peaks.

Figure 5 shows the absorption fiber spectra in the lower wavelength region using fibers 1km in length. The wet silica spectrum is a good example of the principal pattern spectral positions and relative intensities of the OH-bands. Each of the overtones at 724 nm, 943 nm and 1383 nm is accompanied by two less-intense combination modes located on the short wavelength tail of the corresponding overtone absorption band. The basic attenuation is almost entirely governed by Rayleigh scattering. This is the reason why the experiment measures only the seventh overtone $7\nu_3$.

Table 2 shows the measured OH-absorption bands in wet Fused Silica (F100), starting with the fundamental mode ν_3 and going all the way up to the seventh OH-overtone at $7\nu_3 = 444$ nm. The Table also shows combination modes with the SiO_4 tetrahedron vibration ν_1 . We note that a collision between a proton or an oxygen atom of the OH-molecule and a WIMP could excite a number of vibrations detectable by a Bialkali photocathode. We conclude that the following wavelengths fall in the reach of a Bialkali photocathode: 724, 685, 651, 593, 566, 506, and 444 nm.

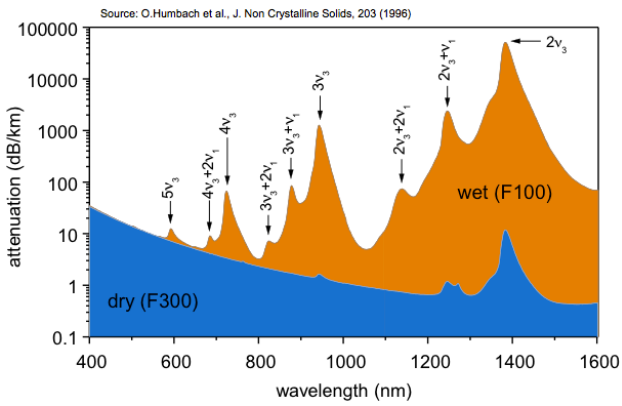


Fig. 5. Attenuation spectra of wet (F100) and dry (F300) silica core fibers, each about 1 km-long. At lower wavelength the attenuation is dominated by Rayleigh scattering, which probably prevents the measurement of even higher overtone modes. The parameter ν_3 corresponds to the OH-molecule fundamental mode, and the parameter ν_1 to the SiO_4 tetrahedron vibration fundamental mode [9].

Table 2. Measured OH absorption bands in wet Fused Silica (F100), starting with the fundamental mode, ν_3 , and going all the way up to the seventh OH-overtone at $7\nu_3 = 444$ nm. The error on the wavelength measurement is ± 1 nm [9].

| OH-band identity | Measured wavelength [nm] |
|-------------------|--------------------------|
| ν_3 | 2722 |
| $\nu_3 + \nu_1$ | 2212 |
| $\nu_3 + 2\nu_1$ | 1894 |
| $2\nu_3$ | 1383 |
| $2\nu_3 + \nu_1$ | 1246 |
| $2\nu_3 + 2\nu_1$ | 1139 |
| $3\nu_3$ | 943 |
| $3\nu_3 + 2\nu_1$ | 825 |
| $4\nu_3$ | 724 |
| $4\nu_3 + \nu_1$ | 685 |
| $4\nu_3 + 2\nu_1$ | 651 |
| $5\nu_3$ | 593 |
| $5\nu_3 + \nu_1$ | 566 |
| $6\nu_3$ | 506 |
| $7\nu_3$ | 444 |

3. OH-molecular vibrations in Pure Water

An obvious question concerns the possibility of OH-absorption peaks in pure water. In first approximation water is remarkably transparent in the visible region. However, a detailed examination reveals the existence of absorption peaks, and that their de-excitations could be detected by a Bialkali photocathode. Figure 6 shows measurements of pure water absorption peaks by the authors listed in Ref. [11]. These absorption peaks are well understood in terms of molecular vibration modes of the water molecule. We conclude that the following wavelengths fall in the reach of a Bialkali photocathode: 742, 662, 605, 550, 514, 473, 449, 418 and 401 nm.

It is clear that the efficiency for Dark Matter to generate light in water is very small, and that it will be necessary to deal with single photons. A large volume of water will be required together with an extensive PMT coverage with single-photoelectron response. Examples of such environments are the BaBar DIRC and the Super-Kamiokande experiment.

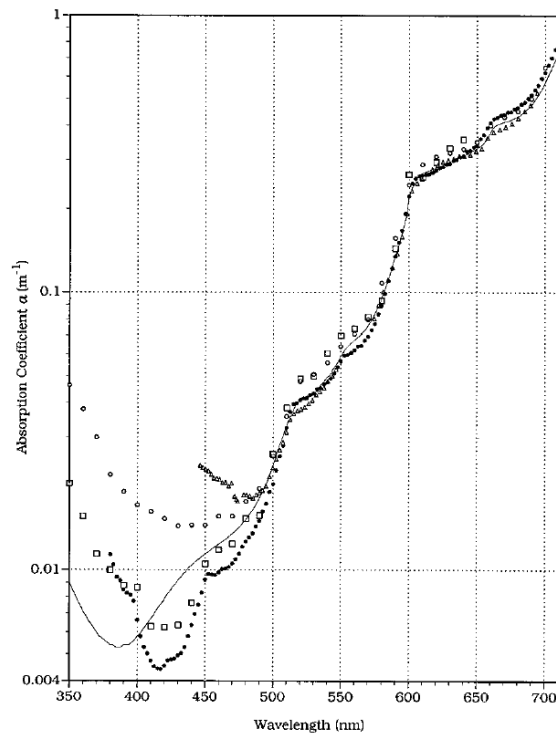


Fig. 6. Comparison of various measurements of absorption in pure water in the visible wavelength region [11].

4. OH-molecular vibrations in NaI(Tl) crystal

Reference [1] used proton recoils from hydrogen atoms to provide a possible new explanation of the DAMA signal. The OH molecular vibrations discussed above represent an additional mechanism which might result in a contribution to this signal.

5. More complicated molecules

In principle, many oils might be used in the study of molecular vibrations. To analyze molecular vibrations in such molecules is much more complicated than to do so in di-atomic molecules. Such molecules have many more modes of vibration. However, if an experiment has a large volume of oil and single-photoelectron sensitivity, it may be sensitive to Dark Matter collisions, even with Bialkali photocathodes. Light emission from such molecules could be calibrated by very low energy neutrons. In response to an earlier version of this article, we have learned, that such an effort is under way in the KamLand experiment.¹

6. Conclusion

This paper provides a scheme within which molecular vibrations might play a role in the discovery of a light WIMP. We have described in detail the vibrations of OH-molecules in Fused Silica and in pure water. A very small energy deposit at the level of 2-3 eV is needed to excite such vibrations, with the corresponding de-excitations detectable by a Bialkali photocathode. This is a factor of 2 lower in energy than is needed to excite the classical scintillation mechanism in a NaI(Tl) crystal. Since the efficiency of the light emission via this process is small, a very large radiator volume of *e.g.* water, Fused Silica or oil is required, together

¹ J. Learned, Univ. of Hawaii, private communication.

with extensive detector coverage with single-photon detection capability.

7. Acknowledgements

I thank Prof. J. Learned for his enthusiastic support of this idea and willingness to search for a DM signal in the KamLand experiment.

References

- [1] J.Va'vra, "A New Possible Way to Explain the DAMA Results", Physics Letters B, 735 (2014) 181.
- [2] R.Bernabei *et al.*, Eur.Phys.J. C 73, 2648 (2013).
- [3] J.Va'vra, "Molecular excitations: a new way to detect Dark matter", arXiv:1402.0466v1, Feb.3, 2014, and arXiv:1401.6995v1, Jan.27, 2014.
- [4] P.M.Morse, Phys. Rev. 34, 57 (1929).
- [5] P.M.Morse, http://en.wikipedia.org/wiki/Morse_potential, 2014.
- [6] K. Gericke, Physical Chemistry lectures, <http://www.pci.tu-bs.de/aggericke/PC4e/>.
- [7] G.P. Smith and D.F. Crosley, J. of Geophysical Research, 95, 16427 (1990).
- [8] Y. Matsumi, *et al.*, Bull. Chem. Soc. Jpn., 75, 711 (2002).
- [9] O.Humbach *et al.*, Journal of Non-Crystalline Solids 203, 19 (1996).
- [10] H.Leutz and G.Schulz, Zeitschrift fur Physik 192, 299 (1966).
- [11] R.M. Pope and E.S. Fry, Appl. Opt. 36, 8710 (1997).